Names, symbols, and relative atomic masses of certain elements

The relative atomic mass (the term atomic weight was previously used) given in the table below is scaled relatively to the isotope ^{12}C taken as 12 exactly and is given to 4 significant figures. Only the elements most commonly encountered in pharmaceutical analysis are included.

Most elements have more than one naturally occurring isotope and the variation in the relative abundance of these isotopes influences the precision with which the relative atomic mass of an element in nature can be quoted. Among the elements included in the table, the relative atomic mass of samples in nature of boron, calcium, lead, strontium and sulfur may differ by more than one in the fourth significant figure.

Name	*Symbol*	*Relative atomic mass*
Aluminium	Al	26.98
Antimony	Sb	121.75*
Arsenic	As	74.92
Barium	Ba	137.3
Bismuth	Bi	209.0
Boron	B	10.81
Bromine	Br	79.90
Cadmium	Cd	112.4
Calcium	Ca	40.08
Carbon	C	12.01
Cerium	Ce	140.1
Chlorine	Cl	35.45
Chromium	Cr	52.00
Cobalt	Co	58.93
Copper	Cu	63.55*
Fluorine	F	19.00
Gold	Au	197.0
Helium	He	4.003
Holmium	Ho	164.9
Hydrogen	H	1.008
Iodine	I	126.9
Iron	Fe	55.85*
Lanthanum	La	138.9
Lead	Pb	207.2
Lithium	Li	6.941*
Magnesium	Mg	24.31
Manganese	Mn	54.94
Mercury	Hg	200.6*
Molybdenum	Mo	95.94
Nickel	Ni	58.71*
Nitrogen	N	14.01
Oxygen	O	16.00*
Phosphorus	P	30.97
Platinum	Pt	195.1*
Potassium	K	39.10*
Ruthenium	Ru	101.1*
Selenium	Se	78.96*
Silicon	Si	28.09*
Silver	Ag	107.9
Sodium	Na	22.99
Strontium	Sr	87.62
Sulfur	S	32.06
Thorium	Th	232.0
Tin	Sn	118.7*
Titanium	Ti	47.90*
Tungsten (Wolfram)	W	183.85*
Uranium	U	238.0
Vanadium	V	50.94*
Zinc	Zn	65.38
Zirconium	Zr	91.22

* Values are considered reliable to ± 1 in the last digit or ± 3 when followed by an asterisk.